# TYPES OF CHEMICAL REACTIONS

# DISPLACEMENT (SINGLE AND DOUBLE)

#### INSTRUCTIONS:

Balance the equations below. Indicate the type of displacement reaction by writing SINGLE or DOUBLE in the space provided. Examples are provided below.

 $1Cu + 2 AgNO_3 \rightarrow 2 Ag + 1 Cu(NO_3)_2$ 

Reaction: SINGLE DISPLACEMENT

1 CuSO<sub>4</sub> + 2 NaOH → 1 Na<sub>2</sub>SO<sub>4</sub> + 1 Cu(OH)<sub>2</sub>

Reaction: DOUBLE DISPLACEMENT

#### QUESTIONS:

1) 
$$\_Cl_2 + \_KI \rightarrow \_KCI + \_I_2$$
 \_\_\_\_\_DISPLACEMENT

2) 
$$K_2CO_3 + BaCl_2 \rightarrow KCl + BaCO_3$$
 DISPLACEMENT

3) 
$$Na + MgCl_2 \rightarrow NaCl + Mg$$
 DISPLACEMENT

4) 
$$\_Al + \_CuCl_2 \rightarrow \_AlCl_3 + \_Cu$$
 \_\_\_\_\_DISPLACEMENT

5) 
$$\_Al + \_Pb(NO_3)_2 \rightarrow \_Al(NO_3)_3 + \_Pb$$
 \_\_\_\_DISPLACEMENT

7) 
$$\_Al_2(SO_4)_3 + \_Ca_3(PO_4)_2 \rightarrow \_AlPO_4 + \_CaSO_4$$
 DISPLACEMENT

8) 
$$\_FeBr_2 + \_K_2CO_3 \rightarrow \_FeCO_3 + \_KBr$$
 DISPLACEMENT

10) \_\_CuS + \_ KCl 
$$\rightarrow$$
 \_ CuCl<sub>2</sub> + \_ K<sub>2</sub>S \_ \_ \_ DISPLACEMENT

# **Double Displacement Reactions Answer Key**

**David E. Goldberg** 

#### **Double Displacement Reactions Answer Key:**

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