

Student Exploration: Stoichiometry (ANSWER KEY)

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Student Exploration: Stoichiometry

Vocabulary: Avogadro's number, balanced equation, cancel, coefficient, dimensional analysis, molar mass, mole, molecular mass, stoichiometry

Prior Knowledge Questions (Do these BEFORE using the Gizmo.)

A 250 mL glass of orange juice contains 22 grams of sugar. How much sugar is in a two-liter (2,000 mL) bottle of orange juice? _____

It requires two sticks of butter to make a batch of 20 cookies. How much butter will it take to make 150 cookies? _____

Gizmo Warm-up

Just as a cook follows a recipe to decide how much of each ingredient to add, a chemist uses **stoichiometry** to determine the amounts of substances involved in chemical reactions. The Stoichiometry Gizmo™ allows you to try your hand at figuring out the amounts of reactants and products that take part in a chemical reaction.

To begin, check that this equation is shown:



Look at the **coefficients** (such as the "3" in 3CO) in front of each substance in the equation. The coefficients tell you how many molecules or atoms take part in a chemical reaction. In the spaces below, list the number of each molecule or atom in the equation:

Fe₂O₃, ____ CO, ____ Fe, ____ CO₂, ____

In a **balanced equation**, the same number of each kind of atom is shown on each side of the equation. Calculate the number of iron (Fe), oxygen (O), and carbon atoms (C).

Reactants Iron: ____ Oxygen: ____ Carbon: ____

Products Iron: ____ Oxygen: ____ Carbon: ____

Based on these values, is the equation balanced? ____

Activity A:	Get the Gizmo ready:	
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Explore Learning Stoichiometry Answer Key

SJ Ball



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